

SAPONIFICATION REACTION OF ETHYL ACETATE WITH SODIUM HYDROXIDE



Leire Unanue, Clara López de Aguilera and Marta Apeztegia

INDEX

1. AIM
2. THEORETICAL BACKGROUND
3. EXPERIMENTAL PROCEDURE
4. BIBLIOGRAPHY

AIM

- Monitor the kinetics of a reaction by looking at the concentration of one of its reactants.
- Check that the reaction is a first order reaction with respect to both reactants.
- Obtain the speed constant at the reaction temperature.

THEORETICAL BACKGROUND



| | | | | | |
|----|-----|-----|---|---|-----|
| 1. | b | a | 0 | 0 | a>b |
| 2. | b-X | a-X | X | X | |

1 is the initial situation, and once time is passed (t), we finish in the second one, in which a quantity of X mol/L of each reagent will have been consumed.

THEORETICAL BACKGROUND

$$\frac{dx}{dt} = k[\text{NaOH}]^{\alpha} [\text{EtAz}]^{\beta} = k(a-x)^{\alpha} (b-x)^{\beta}$$

Equation 1: Kinetic equation

$$k \cdot t = \frac{1}{(a-b)} \ln \frac{b(a-x)}{a(b-x)}$$

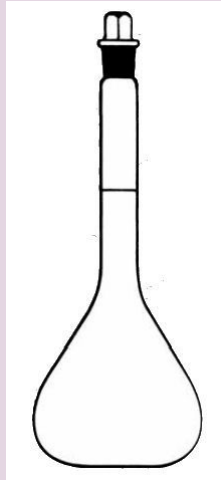
Equation 2

$$\ln \left(\frac{V_0 - V_t}{V_{\infty} - V_t} \right) + \ln \frac{b}{a} = k(a-b) \cdot t$$

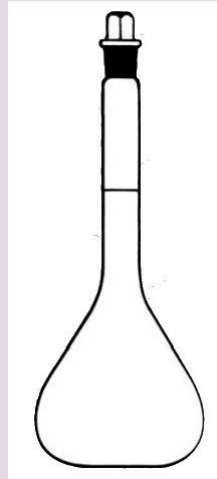
Equation 3

EXPERIMENTAL PROCEDURE

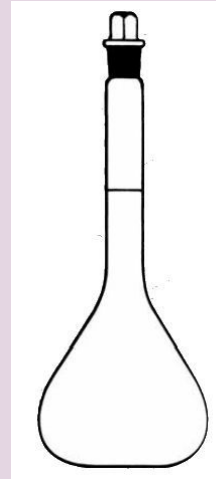
1. Initial dissolutions



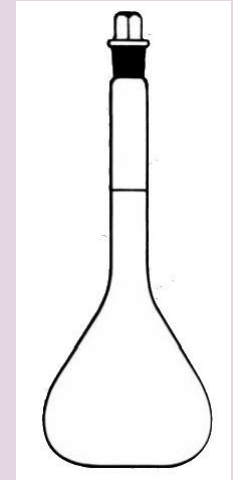
1L NaOH
0,025M



250 mL HCL
0,05M



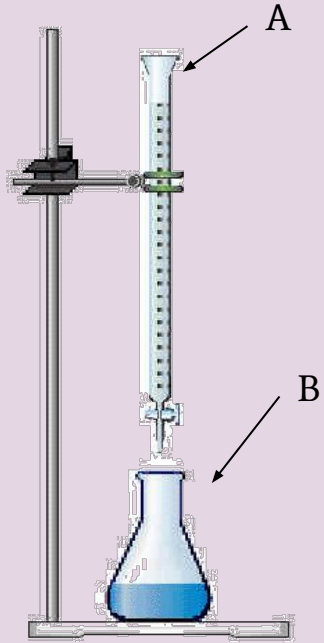
KHFt
0,025M



1L AcEt
(1,6 mL solute/1L
dissolution)

EXPERIMENTAL PROCEDURE

2. Titration



| TITRATION | 1 | 2 |
|---------------------|------|------|
| A: Reagent solution | KHFt | NaOh |
| B: Analyte | NaOh | Hcl |

EXPERIMENTAL PROCEDURE

3. Heat the dissolutions in an isothermal bath

ISOTHERMAL BATH

Until you reach 25°C

15mins aprox.



100mL NaOH



100 mL AcEt

EXPERIMENTAL PROCEDURE

4. HCl



x6

10mL Hcl

EXPERIMENTAL PROCEDURE

5. Reaction NaOH and Acet



100mL NaOH

Quickly!



100mL AcEt

Leave it for 3
minutes

*Keep the reaction Tcte

EXPERIMENTAL PROCEDURE

6. Pour the reacting mixture with HCl



100mL NaOH
+
100mL AcEt

25ml



10mL HCL

Repeat this operation after.

- 7min
- 13min
- 20min
- 30min
- 45min

(Remember to back-titrate the excess acid)

After the last extraction, cover the erlenmeyer and leave it to react at room T until the following day.

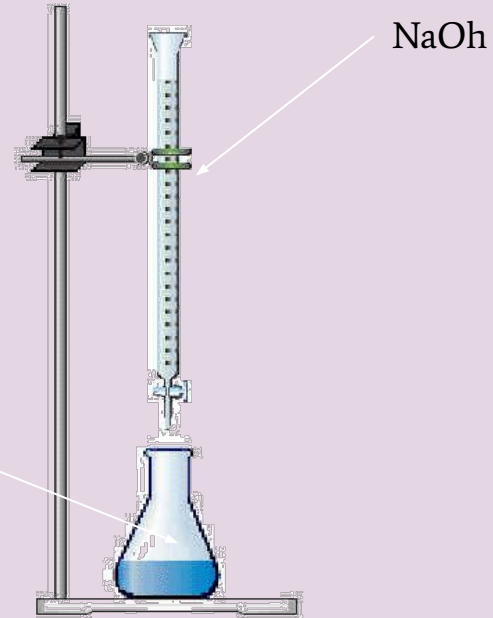
EXPERIMENTAL PROCEDURE

7. Titrate



Reacting mixture

25mL



BIBLIOGRAPHY

https://egela.ehu.eus/pluginfile.php/2546939/mod_resource/content/1/9_SaponificationRxn.pdf