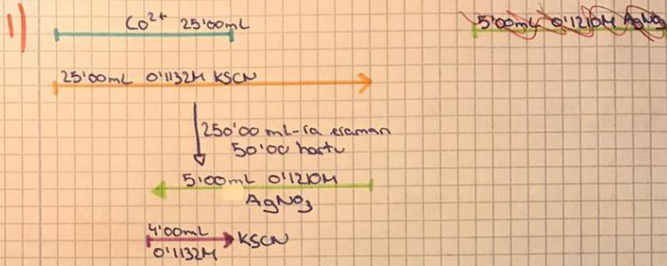


Kimika analitika : 2. partziala

Ioritz Arburua Leonis



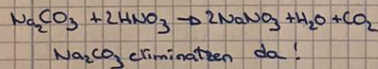
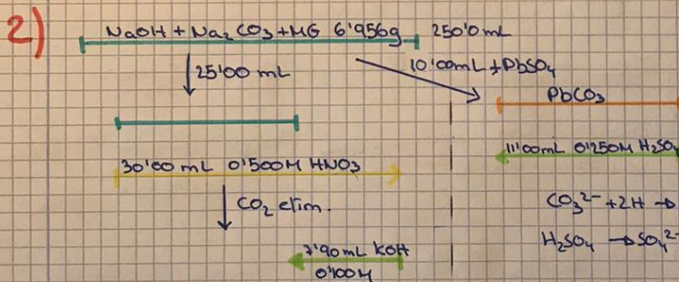
$$n_{\text{SCN}^-} = n_{\text{Ag}^+} - n_{\text{SCN}^- \text{ gehiegia}} = 5.00 \cdot 0.1210 - 4.00 \cdot 0.1132$$

$$= 0.152 \text{ mmol} \xrightarrow{\frac{25.00}{50.00}} 0.7610 \text{ mmol disoluzioa } 25.00 \text{ mL}$$

$$n_{\text{Co}^{2+}} = \frac{1}{2}(n_{\text{SCN}^-} - n_{\text{SCN}^- \text{ gehiegia}})$$

$$= \frac{1}{2} 25.00 \cdot 0.1132 - 0.761 = 1.034 \text{ mmol}$$

$$[\text{Co}^{2+}] = \frac{1.034}{25.00} = 0.04136 \text{ M}$$



$$n_{\text{NaOH}} = n_{\text{HNO}_3} - n_{\text{KOH}}$$

$$= 30.00 \cdot 0.500 - 3.90 \cdot 0.100$$

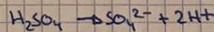
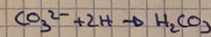
$$n_{\text{NaOH}} = 14.21 \text{ mmol (25 mL)}$$

$$142.11 \text{ mmol (250 mL)}$$

$$m_{\text{NaOH}} = 142.11 \cdot 39.997$$

$$= 5683.6 \text{ mg}$$

$$\% \text{NaOH} = \frac{5683.6}{6956} \cdot 100 = \%81.71$$



$$n_{\text{CO}_3^{2-}} = n_{\text{H}_2\text{SO}_4} = 11.00 \cdot 0.250$$

$$= 2.75 \text{ mmol} \rightarrow 10.00 \text{ mL}$$

$$68.75 \text{ mmol} \rightarrow 250.0 \text{ mL}$$

$$m_{\text{Na}_2\text{CO}_3} = 68.75 \cdot 105.99 = 7286.81 \text{ mg}$$

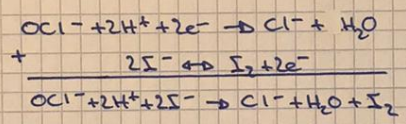
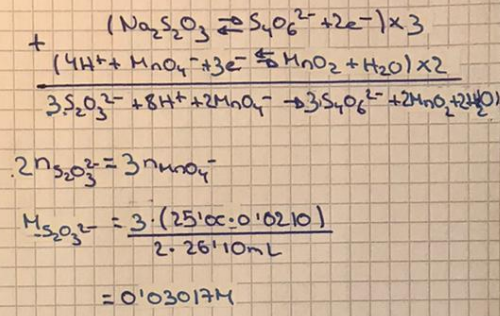
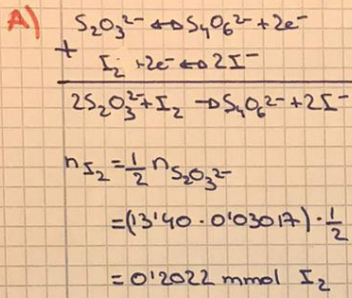
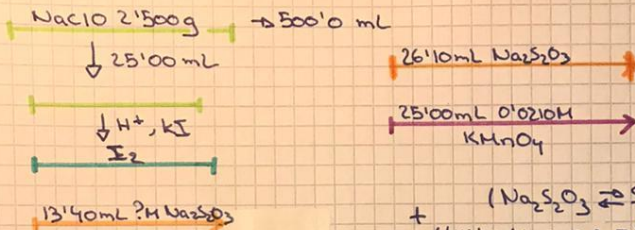
$$7.28681 \text{ g} \cdot 6.956 \text{ g}$$

Wontait akatsen bat egin dut lajin osak baino pisu gehiago daukalako

Kalkuluak behin eta berriro errepasatu ditut beti emaitza bera

$$\% \text{Na}_2\text{CO}_3 = \frac{7286.81}{6956} \cdot 100 = \%104.76$$

3.



$n_{I_2} = n_{OCl^-} = 0'2022 \text{ mmol (25 mL-tan)}$
 $= 4'043 \text{ mmol (500 mL-tan)}$
 $m_{NaClO} = 4'043 \cdot 74'4 = 300'81 \text{ mg} \rightarrow \% = \frac{300'81 \cdot 100}{2500} = \%12'03$

B) $\frac{300'81 \text{ mg}}{500'0 \text{ mL}} = \%6'016 \gg \%0'2 \rightarrow$ Ezingoa litzateke kuxiba hari erabili desinfekzioarako

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